

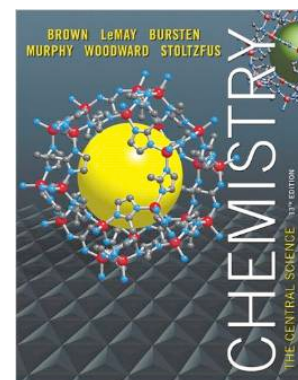
GENERAL CHEMISTRY B

Instructor: Willetta Greene-Johnson, Ph. D., Room 307 Cudahy Science 773-508-3537
Who am I: A chemical physicist (Stanford University/ University of Chicago) trained in spectroscopy and statistical mechanics. I'm interested in (1) thermodynamics (2) swarming (3) entropy (4) producer, pianist, composer, orchestrator, sequencer, and conductor. I'm a classically trained pianist. My vocal ensemble has recorded several projects—hopefully can do an EP later this year. One of my songs was doubly tracked on a Grammy award winning vocal CD in 2004. The same song was recorded on DVD (released April 2008). That song has been recorded by five different artists. I'm also getting more orchestral arrangement contracts—including the Memphis Symphony Orchestra. Orchestras rock!

Physical Office Hours: **Wednesday 10:30 A – 11:20 A CS-307**
 Email Office Hours (ONLY): **Thursday 10:00 A – 11:00 A wgreene@luc.edu**

Required:

1. Required: Chemistry, the Central Science. 13th ed. Theodore L. Brown, et. al. Boston: Pearson Prentice Hall: 2014
 ISBN: 978-0321910417. The *Mastering Chemistry* asset is NOT required in my section
2. **Chemistry 102 Course Packet**, authored by the instructor. This essential lecture packet is available online at www.universityreaders.com. The course packet will be mailed to you within a few days of ordering, but you'll have immediate online access to 20% or so once order is completed.



Meetings: Lectures are scheduled MWF in FH-133, at 11:30 P – 12:20 P. You must also be registered in discussion section 009 - 011.

Discussions: meet on Fridays according to the following schedule:

Section	Instructor	Location	Time
009	Dr. Greene-Johnson	FH-007	F: 12:35 P – 1:25 P
010	Dr. Greene-Johnson	FH-007	F: 1:40 P – 2:30 P
011	Dr. Angela Mahaffey	FH-007	F: 2:45 P – 3:35 P



Due to the large number of students / sections that are matriculated through this course yearly, there can be absolutely no alteration of this schedule.

Course Description: A study of chemical principles with emphasis on the development of a scientific attitude and an understanding of the fundamental concepts of chemistry. Robust pre-calculus concepts will be particularly emphasized the first third of the semester.

Calculators: Any scientific calculator is sufficient, however calculators cannot be shared while exams are in progress and their cases/covers must be removed. Be sure that you are familiar with **your calculator** and that its batteries are **in good condition**, especially on exam days. The student is responsible for remembering to bring his calculator on an exam day.

SAKAI Connection: The syllabus, homework assignments for the semester, discussions, and discussion answers will be posted at the following website: www.luc.edu, look under *Technology*, click on **Sakai**). Students possessing a Loyola email address are able to access this site.

Objective of this course in grander detail:

By the conclusion of this course, the student should experience the following outcomes:

1. **Understand the fundamental principles of physical chemistry**
2. **Acquire a knowledge base of basic terminology and classifications**
3. **Apply concepts creatively as well as methodically to solve multi-tiered problems**
4. **Know how to rank, estimate, analyze and critically evaluate a range of models**
5. **Gain a broader understanding of the role of chemistry in human endeavor**
6. **Appreciate the collaborative and global effort of the scientific enterprise**

Specifically the engaged student should improve in her or his ability to


- **Comprehend the following concepts:**
 - Kinetics, reaction rate, Arrhenius equation, rxn mechanism, rate limiting step
 - Dynamic equilibrium and Equilibrium constants
 - Titrations, Buffers, pH, Lewis and Arrhenius A/B models, Solubility product
 - Complex ions and octahedral Crystal Field theory
 - Electrochemical cells (batteries, fuel cells, transduction, respiration)
 - Spontaneity vs. nonspontaneity, entropy, cyclic processes, and free energy
 - Introduction to nuclear chemistry
- **Identify reagents and general chemical processes:**
 - Identify acids, bases, acidic and basic salts, buffers, solubility rules (Chemistry101)
 - write appropriate net ionic prototypical rxns in aqueous solution
 - be able to closely estimate and accurately calculate pH
 - equilibrium constant K and how it predicts spontaneity status of a reaction
 - complex ion, Lewis bases, geometry of “simpler” complex ions and compounds.
- **Assess outcome feasibility:** estimate energy and entropy of chemical processes
- **Work and exchange ideas with others:** cordially solve weekly group problems
- **Appreciate the impact of chemistry:** realize better how chemistry impacts life processes, technology, local, and global issues.
- **Contribute constructively:** as a science-literate, ethically responsible citizen and voter.





Later this semester, you will receive an emailed invitation to assess me via the **IDEA** (Individual Development and Educational Assessment). The form provides a thorough diagnostic of how successfully students think the instructor realized the objectives boxed above, as well as the value of the course and other contextual experiences. This opportunity will be available online at <http://www.luc.edu/IDEA> for a one-week time window only,

IDEA manual states: 'As student raters, please be aware that the results of your ratings for this class will be included as part of the information used to make decisions about promotion/tenure/salary increases for this instructor.

Fairness to both the individual and the institution require *accurate and honest* answers.'

Chemistry 102 Schedule of Topics

Date	Day	Topic	Chapter	Pages (approx)
1/12 JAN	M	Rate of Reactions	14	574-582
1/14	W	Integrated Rate Law	14	582-590
1/16	F	Arrhenius Equation	14	591- 600
1/21	W	Rxn mechanisms; Rate limiting step	14	601-606 Friday=1st discussion
1/23	F	Catalysis	14	607-614
1/26	M	Gas phase Equil'm	15	628-637
1/28	W	Const. K; Equil. Table	15	632-644; 644-650
1/30	F	van't Hoff equation	15	648-656
2/2 FEB.	M	Le Chateliér's Principle	16	657-660
2/4	W	Acid/base categories	16	670-679, 684-687
2/6	F	pH, K_a and K_b	16	680-684; 688-695
2/9	M	Weak bases, Conjugates, Salt Solutions	16	696-699, 705-712
2/11	W	Review 14-16 25 min		OPTIONAL
2/13 	F	Exam 1: Ch.14-16	IN CLASS	Seating: SKIP Every 3rd row!
2/16	M	Binary Acids/ strength	16	705-712
2/18	W	Buffers	17	729-737
2/20	F	Titrations SA/SB;SA/WB;SB/WA	17	738-747
2/23	M	K_{sp} and ppt equil'm	17	748-751
2/25	W	Common Ion Effect	17	726-728, 751-753
2/27	F	Complex Ions (Lewis Acid/Base)	23 17	1002-1006 756-758
3/2-6 MAR	M-F	SPRING BREAK	sPrInG BrEaK	Spr.Break Worksheet
3/9	M	Ligands, Nomenclature	23	1007-1016
3/11	W	Crystal Field Theory	23	1020-1026, 1030-3
3/13	F	Review: 16,17,23 25 min		Optional no handout
3/16	M	EXAM 2: Ch.16,17,23	IN CLASS	Seating: SKIP Every 3rd row!
3/18	W	Spont. And Temp. Statistical origin (physics)	19	812-818 821-828 (opt)
3/20	F	Entropy, standard, 2 nd Law	19	818-821,828-831
3/23	M	Gibb's Energy and Temp Gibbs Energy and K	19	831-838 838-841

Date	Day	Topic	Chapter	Pages
MAR. 24 (5 PM) LAST DAY TO WITHDRAW WITH A GRADE OF W				
3/25	W	Coupled Reactions Hess's Law for ΔG , ΔS	19	842-844
3/27	F	Electrochemistry Redox Review Voltaic Cell $E^\circ = E_{\text{ox}}^\circ + E_{\text{red}}^\circ$	20	856-865 858-868
3/29	M	E_{red}° , E_{ox}° , Spontaneity	20	868-874
3/31	W	Work and ΔG	20	876-879
4/2-4/6 	Th-M 	EASTER BREAK 		
4/8	W	Graphical Technique	20	In class only
4/10	F	Nernst Equation E° and equil'm const. Applications; Electrolysis	20	880-886 886-892 893-896
4/13	M	Nuclear Rxns; Stability; 1 st order Kinetics, Geological Dating	21	908-919 920-928
4/15	W	REVIEW 19-21 (25 min)		Optional may not be handout
4/17	F	Exam 3: Ch. 19-21 (what we get to in Ch. 21)	IN CLASS	Seating: SKIP Every 3rd row
4/20	M	Transmutation	21	918-920
4/22	W	Einstein: mass/energy Fission/fusion	21	921- 931 932-938
4/24	F	REVIEW for Final (minimally 25 min)		In lecture may continue in disc'n
4/27 APRIL	M	FINAL EXAM : Ch. 14-17,19-21, 23	1:00 P - 3:00 P	TBA (FH-133 probably)

Representative Problems, End of Chapter Problems & Discussions:

Students who are making good progress in the course should be able to solve, independently, most or all of the end-of-chapter problems in the text. You should attempt to work out as many example problems and end-of-chapter problems as possible before taking exams. A group of representative problems is listed below as assigned problems. The solution manual with the worked out problems will be kept on reserve in Cudahy Library. A comprehensive review containing additional problems will be posted approximately one week before the midterm exams, which also serves for the final exam.

Discussions count 10% of grade, and should be attended. Discussion problems will be given to be attempted by groups of 3-4 students within the 50-minute discussion. Students must stay the entire period (unless otherwise instructed) and work on assigned discussion problem(s) to earn up to 10% of grade. The solutions will be posted on blackboard, discussions will be returned by the following discussion, or during the same week if a Friday exam occurs in that week. The student is strongly encouraged to attempt all suggested problems (text-book and discussion) and contribute significantly to the group discussion activity. **Student's extent of group participation will be noted** and mentioned in any recommendation letters.

Assigned Exercises:

chapter	page	Problems (*means more involved;** means unassigned reading may be required)
14	615:	3a, 5, 7, 9, 12, 19, 21 a - c, 23, 25 (sim. to 23), 27, 29, 33, 37, 38, 41, 43, 45, 50 (glucose is C ₆ H ₁₂ O ₆ !) 57, 61, 63*, 67, 73, 75, 91, 95, 99, 105, **117.
15	661:	1, 2, 6, 9, 13, 15, 17, 23, 25, 27, 28, 33a, 35, 37, 39, 45 find [Cl ₂] _{eq} , next PV = nRT 51, 53, 55, 57, 61, 63, 65, 67, *74, 79, *80. I don't overly emphasize K _c vs K _p but for practice, do #24, use formula 15.14 on page 636 % ionization = $\frac{x}{x_0} \times 100\%$; x ₀ = starting concentration, unit usually M;
16	715:	1, 3, 10, 15, 19, 21, 23, 27, 29, 33, 37, 43, 45, 47, 53, 55, 57, 61, 65, 69, 71, 73, 77, 79, 81, 85, 109, 116*, 119*; 1**, 12**, 95**, 99** **compares various A/B models
17	766:	1, 6, 10, 5, 17 (% ionization ↔ [H ⁺]), 19, 21, 27, 29, 33, 35, 37, 41, 43, 47, 53, 55, 57, 61, 69, 73, 83; 12**, 67** note: K _a , K _b and K _{sp} may be involved in some problems concurrently.
19	845:	1, 3, 7, 9, 11, 13, 15, 25, 35, 37, 41, 43, 49, 53, 59, 61, 65, 69, 71, 79b, 83, 85, 98* 1 st find ΔG [∘] (appendix), 2 nd K: use ΔG [∘] = -RTLnK, 3 rd : ΔG = ΔG [∘] + RTLnQ ; 10*, 75*, 94*, 106**
20	899:	4, 6, 7, 12, 13, 15, 19, [23, 25 review balance redox rxn] 29a-e, 31, 37, 39, 43, 53, 55, 57, 59, 61a, 65, 67, 99, 100; 103*, 87**, 73*, 75* (these are longer than usual)
21	946	1, 5, 9, 11, 13, 17, 21, 29, 35, 37, 47, 49, 55, 61 extra: 72, 74, 63*, 80*, 31**
23	1031:	15, 16, 23, 25, 27, 35, 37, 41, 43, 47, 55, 57, 59, 61, 63; extra nerdy: 73, 74, 80

Academic Honesty:

All students are responsible for exercising the highest level of academic honesty while taking exams. They should peruse the College of Arts & Science policy on plagiarism/cheating, stated at:

<http://www.luc.edu/media/lucedu/cas/pdfs/academicintegrity.pdf>

As in the past, cheating will be SEVERELY dealt with, *minimally* costing the offender a grade of “zero” for the item that was submitted and this grade cannot be dropped. Additionally, the incident must be reported to the Chemistry Department Chair and the Office of the CAS Dean. Depending on the seriousness of the incident, additional sanctions may be imposed. Which has happened before.

Grading Scheme:

There are two grading schemes, and whichever one yields the higher grade will be employed after the final has been taken:

If all midterms went fairly well: 20 % midterm, 10% discussion, 30% final

If one midterm not so good: 20% the other two, 10% discussion, 50% final

Grading Scale:

Grade Scale:	A	≥ 91	A-	88-90	
B+	85-87	B	81-84	B-	78-80
C+	75-77	C	71-74	C-	67-70
D+	64-66	D	61-63	F	< 60

Examinations

Three hour exams and the final exam will be given on **Feb. 13, Mar. 16, April 17, and April 27**, respectively, also noted in the schedule. 90% of your course grade will be determined from these as explained further below. The other 10% will be determined from your discussion grades. The exams are cumulative, *i. e.*, may include material that has been queried on previous exams. The final exam is comprehensive and cumulative.

Final Exam:

Please note that **the final examination must be taken**. Failure to take the final exam will result in the grade “F”. If a student has missed the final for some valid reason, she/he must present the Dean's office with reasonable proof of illness or accident, verified by a doctor's note, police report, *etc.*, in order to take the makeup final on a single date designated by the Dean's office. A fee may apply and ***there is no guarantee that content on the make-up final will correspond to / have the same weight as the scheduled final.***

Laboratory:

Chemistry 112, the general chemistry laboratory course, should be taken concurrently with the lecture course in general chemistry. The lecture and the laboratory courses are graded independently. Students should first consult the Chemistry Department Bulletin opposite the wall facing the chemistry office for information. Then, if they still have unresolved issues, they should contact Dr. Angela Boerger, the administrator of the laboratories.

Flanner Hall-133 Room Instructions on Exam Days

- 1) When you enter the auditorium, **go to the front** and **place your book bag there**. **Remove your calculator slipcover** and placed it in book bag.
- 2) Starting from the first row nearest the lectern, **sit quickly** in every other seat and **skip every third row**. This vacant third row provides an aisle for the proctor to walk through and address any appropriate questions that student may have during exam. Do not try to sit with friends or near one's usual area. The exam is only **50 minutes**, so excessive delays will cut into exam-taking times.
- 3) **Place your student ID conspicuously** on your desk so that attendance may be noted (during exam).
- 4) Have **several pencils/pens, eraser**, etc. and a **calculator** in good **working** order.
- 5) Proctors have been instructed to **confiscate the exams of any student** using a calculator with its slipcover in place or **whose actions are suspect**.
- 6) **Read over the entire exam**. You may find a problem in the middle, or at the end, that suits you better to start. The three or so minutes spent glancing over the entire exam will be more than compensated for by the strategy and priorities that you formulate. The recommended **order to do problems** is:
 - (1) what you **know well FIRST**
 - (2) what you're sure you **can at least start NEXT**
 - (3) what you **haven't have a clue LAST**

I have tried to arrange problems in a reasonable order, but my perception and the student's will certainly differ in some aspects. So, take a few minutes to read over the exam and **devise your own strategy**.
- 7) When you have **concluded, turn in your exam** to a proctor. Then **leave as quietly** and as expeditiously as possible as to not disturb other exam takers.
- 8) Normally exams administered on Friday will be returned no later than the following Wednesday.

Advanced Studies Recommendation Protocol

Later in your student career, you may require a letter of recommendation (LOR) for graduate school, medical school, or the like. If I am chosen among your recommenders, the following policy ensues:

1. **Deadline for LOR (letter of recommendation) requests: Jan 15th of the application year.**
2. Student must generally possess GPA of 3.5 or above. However, if my time allows, a student might be considered if she or he presents a **written explanation** that reveals an exceptional circumstance accounting for a lower grade point average.
3. Student must provide attached in one email, a document listing his/her correct GPA, contact information, deadline(s), and also all chemistry, biology and physics courses and labs that the student has take—in the following format (or Committee format, if you are applying through committee):
 - a. **GPA**
 - b. reliable, current email and telephone # that student checks *regularly*
 - c. **DEADLINE**
 - d. Table with header: course taken, instructor, grade

Example:

Course	Semester/year	Instructor	Grade
Chemistry 101	Fall /2014	Dr. WGJ	A-
Biology 210	Spring /2015	Dr. Barbara Haas	A

- e. If applying through Committee, you **MUST** include **PDF signed waiver** (2 pgs) in email with other items.
 - f. If applying “outside the Committee”—see items 5, 6 below, a list of all schools of the applicant and **ALL of their DEADLINES**.
 - g. All cover forms, application packages, envelopes in one binder, folder, or otherwise secure containment, with like items paper-clipped together.
4. I'd love to read your personal statements, even in rough draft form. It tells me something about you and helps me to shape a recommendation. This article is not required, but I recommend it.
 5. **It is STRONGLY recommended that the student apply through the Loyola Pre-Health Advisory Committee.** Well-regarded by the medical /dental/ pharmaceutical community, the Committee's voice of endorsement will increase the merit of the student's application. Their method also assures that the student's personal statement is strong and well written. If the student applies via Committee, I will be contacted.
 6. **APPLICATIONS OUTSIDE COMMITTEE:** If a student who I can recommend elects to apply apart from the Pre-Health Advisory committee, she/he must perform steps 2-4 and email me at wgreene@luc.edu. Online LOR uploading protocols (AMCAS, PTCAS, Interfolio, *etc.*) are **STRONGLY** preferred.

Because of the number of requests, your LOR won't be started until all items in 2 are fulfilled

LOYOLA UNIVERSITY CHICAGO SPRING CALENDAR 2015

January 11 (midnight)	Sunday	Open registration ends
January 12	Monday	Spring Semester begins Late and Change Registration begins Late registration fees apply
January 19	Monday	Martin Luther King, Jr., Holiday: No classes
January 20	Monday	Late and change registration ends Last day to withdraw without a "W" grade
January 25	Sunday	Last day to drop class(es) with a Bursar credit of 100%
January 26	Monday	Last day to convert from credit to audit or vice versa
February 8	Sunday	Last day to drop class(es) with a Bursar credit of 50%
February 9	Monday	Summer Registration begins
February 18	Wednesday	Ash Wednesday: Classes meet; Special services
February 15	Sunday	Last day to drop class(es) with a Bursar credit of 20% (zero credit thereafter)
February 23	Monday	Last day for students to submit assignments to change an "I" grade to a letter grade for Fall Semester 2008; Faculty may set earlier deadlines with students
February 23	Monday	Early Alert process begins
March 1	Saturday	Last day to file applications with Deans' offices for degrees awarded in December for this year.
March 2 - 7	Monday-Saturday	Spring Break: No classes
March 9	Monday	Classes resume after Spring Break
March 23 (5:00 P)	Monday	Last day to withdraw with a grade of "W" After this date, the penalty grade of "WF" will be assigned
April 2 – April 6	Th(4:15) – M(4:15)	Easter Holiday
April 13	Monday	Fall Semester Registration begins
April 24	Friday	Spring Semester classes end
April 29	Wednesday	Study Day: No classes
April 27 – May 2	Monday-Saturday	102 Chemistry Final convenes April 27 2015 (Monday) from 1:00 P to 3:00 P

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Manager: Dionne Damico **Email Address:** luc-lsc@bkst.com